Supporting Information for

# Engineering Spin States of Isolated Copper Species in a Metal–Organic Framework Improves Urea Electrosynthesis

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# **Supplementary Figures and Tables**

**Fig. S1** SEM of image Cu<sup>III</sup>-HHTP



Fig. S2 a SEM and b TEM image and EDS elemental mapping of as-made Cu<sup>II</sup>-HHTP



Fig. S3 FT-IR spectra of  $Cu^{II}$ -HHTP and  $Cu^{III}$ -HHTP



**Fig. S4 a** XPS survey curve of Cu<sup>III</sup>-HHTP, **b** Cu 2p spectrum of Cu<sup>II</sup>-HHTP and **c** O 1s spectra of Cu<sup>III</sup>-HHTP and Cu<sup>II</sup>-HHTP

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![](_page_2_Figure_1.jpeg)

Fig. S5 XRD patterns of two catalysts

![](_page_2_Figure_3.jpeg)

Fig. S6 Raman spectra of two catalysts

![](_page_2_Picture_5.jpeg)

Fig. S7 HAADF-STEM image of Cu<sup>II</sup>-HHTP

![](_page_3_Figure_1.jpeg)

Fig. S8 Temperature-dependent susceptibility  $\chi$  for Cu<sup>II</sup>-HHTP and Cu<sup>III</sup>-HHTP catalysts

![](_page_3_Picture_3.jpeg)

**Fig. S9** The optical photograph of the H-type cell and gas purification unit for urea electrosynthesis testing

![](_page_3_Figure_5.jpeg)

Fig. S10 The schematic diagram of *in situ* ATR-FTIR tests

![](_page_4_Figure_1.jpeg)

**Fig. S11 a** UV-vis curves and **b** concentration-absorbance of urea solution with a series of standard concentration (0.0-1.7  $\mu$ g mL<sup>-1</sup>) in 0.1 M KHCO<sub>3</sub> solution. The standard curve shown good linear relation of absorbance with urea concentration (y = 0.1064x + 0.0187, R<sup>2</sup>=0.999)

![](_page_4_Figure_3.jpeg)

**Fig. S12 a** UV-vis curves and **b** concentration-absorbance of NH<sub>4</sub>Cl solution with a series of standard concentration (0.0-3.0  $\mu$ g mL<sup>-1</sup>) in 0.1 M KHCO<sub>3</sub>. b) The standard curve shown good linear relation of absorbance with NH<sub>4</sub>Cl concentration (y=0.0841x+0.04694, R<sup>2</sup>=0.994)

![](_page_4_Figure_5.jpeg)

**Fig. S13 a** UV-vis curves and **b** concentration-absorbance of KNO<sub>2</sub> solution with a series of standard concentration (0.0-6.2  $\mu$ g mL<sup>-1</sup>) in 0.1 M KHCO<sub>3</sub>. The standard curve shown good linear relation of absorbance with KNO<sub>2</sub> concentration (y = 0.1012x + 0.0269, R<sup>2</sup>=0.999)

![](_page_5_Figure_1.jpeg)

**Fig. S14 a** UV-vis curves and **b** concentration-absorbance of N<sub>2</sub>H<sub>4</sub> solution with a series of standard concentration (0.0-6.0  $\mu$ g mL<sup>-1</sup>) in 0.1 M KHCO<sub>3</sub>. The standard curve shown good linear relation of absorbance with N<sub>2</sub>H<sub>4</sub> concentration (y = 0.3694x + 0.1006, R<sup>2</sup>=0.998)

![](_page_5_Figure_3.jpeg)

**Fig. S15 a** UV-vis curves and **b** concentration-absorbance of KNO<sub>3</sub> solution with a series of standard concentration (0.0-6.0  $\mu$ g mL<sup>-1</sup>) in 0.1 M KHCO<sub>3</sub>. The absorbance was measured by UV-vis spectrophotometer. The standard curve shown good linear relation of absorbance with KNO<sub>3</sub> concentration (y = 0.0337x + 0.00147, R<sup>2</sup>=0.999)

![](_page_5_Figure_5.jpeg)

**Fig. S16** Absorption spectra of the electrolyte after  $Cu^{III}$ -HHTP catalysts at various potentials for 2h in N<sub>2</sub> and CO<sub>2</sub>-saturated 0.1 M KHCO<sub>3</sub> solution to quantify NO<sub>2</sub><sup>-</sup>

![](_page_6_Figure_1.jpeg)

Fig. S17 Absorption spectra of the electrolyte after  $Cu^{III}$ -HHTP catalysts at various potentials for 2h in N<sub>2</sub> and CO<sub>2</sub>-saturated 0.1 M KHCO<sub>3</sub> solution to quantify N<sub>2</sub>H<sub>4</sub>

![](_page_6_Figure_3.jpeg)

**Fig. S18** Absorption spectra of the electrolyte after  $Cu^{III}$ -HHTP catalysts at various potentials for 2h in N<sub>2</sub> and CO<sub>2</sub>-saturated 0.1 M KHCO<sub>3</sub> solution to quantify NO<sub>3</sub><sup>-</sup>

![](_page_6_Figure_5.jpeg)

Fig. S19 The linear sweep voltammetry (LSV) of a  $Cu^{II}$ -HHTP and b  $Cu^{II}$ -HHTP catalysts in CO<sub>2</sub>, N<sub>2</sub> and CO<sub>2</sub> + N<sub>2</sub> saturated electrolyte

![](_page_7_Figure_1.jpeg)

**Fig. S20** The urea yield rate and Faradaic efficiencies of Cu<sup>III</sup>-HHTP catalyst at the different potentials

![](_page_7_Figure_3.jpeg)

**Fig. S21** The urea yield rate, Faradaic efficiencies and the corresponding product distribution of  $H_2$  (orange), CO (blue), NH<sub>3</sub> (green), and urea (pink) with N<sub>2</sub> and CO<sub>2</sub> as the feeding gas at various potentials for Cu<sup>III</sup>-HHTP catalysts

![](_page_7_Figure_5.jpeg)

Fig. S22 The chronoamperometric curves  $Cu^{III}$ -HHTP catalysts at various potentials for 2h in N<sub>2</sub> and CO<sub>2</sub>-saturated 0.1 M KHCO<sub>3</sub> solution

![](_page_8_Figure_1.jpeg)

Fig. S23 The chronoamperometric curves of Cu<sup>III</sup>-HHTP catalyst at -0.6 V vs. RHE for 18 h in  $N_2$  + CO<sub>2</sub>-saturated 0.1 M KHCO<sub>3</sub> solution

![](_page_8_Figure_3.jpeg)

**Fig. S24** The Faradaic efficiency and urea production rate of Cu<sup>III</sup>-HHTP catalysts at -0.6 V vs. RHE during five recycling tests

![](_page_8_Figure_5.jpeg)

**Fig. S25** Nyquist plots of electrochemical impedance spectra (EIS) of  $Cu^{II}$ -HHTP and  $Cu^{III}$ -HHTP

![](_page_9_Figure_1.jpeg)

**Fig. S26** CV curves of **a** Cu<sup>II</sup>-HHTP, **b** Cu<sup>III</sup>-HHTP under different scan rates from 4 to 20 mV  $s^{-1}$ 

The electrochemically active surface area (ECSA) can be determined through the double layer capacitance (C<sub>dl</sub>) varying the scan rate of cyclic voltammetry (CV) curves, which is an essential parameter for the evaluation of electrochemical reactivity. The double layer capacitance (Cdl) of the two MOFs samples was determined by the slope of the linear fit of  $\Delta J = Ja - Jc$  at 0.05 V (vs. RHE) and the scan rate (**Fig. S26**). The specific capacitance of a generally slick flat surface is between 20 and 60  $\mu$ F cm<sup>-2</sup>, thus 40  $\mu$ F cm<sup>-2</sup> is used as a reference in this work to calculate the ECSA. The ECSA of Cu<sup>II</sup>-HHTP, (b) Cu<sup>III</sup>-HHTP is 337.5 cm<sup>-2</sup> 570.0 cm<sup>-2</sup>, respectively.

![](_page_9_Figure_4.jpeg)

![](_page_9_Figure_5.jpeg)

In the calculation of the electrochemically active surface area for the three samples below, we have assumed a specific capacitance of 40  $\mu$ F cm<sup>-2</sup> for a flat surface, which is calculated as follows:

$$A_{ECSA}^{Cu^{II}-HHTP} = \frac{13.5 \text{ mF cm}^{-2}}{40 \text{ }\mu\text{F cm}^{-2} \text{ per cm}_{ECSA}^{2}} = 337.5 \text{ cm}_{ECSA}^{2}$$
$$A_{ECSA}^{Cu^{III}-HHTP} = \frac{22.8 \text{ mF cm}^{-2}}{40 \text{ }\mu\text{F cm}^{-2} \text{ per cm}_{ECSA}^{2}} = 570.0 \text{ cm}_{ECSA}^{2}$$

![](_page_10_Figure_1.jpeg)

**Fig. S28 a** The calibration curves for  ${}^{15}NH_2CO{}^{15}NH_2$  solution at concentrations of 0.3-1.2  $\mu$ g/mL, **b** The corresponding calibration curve for  ${}^{15}NH_2CO{}^{15}NH_2$  solution

![](_page_10_Figure_3.jpeg)

**Fig. S29** The urea yield of Cu<sup>III</sup>-HHTP catalyst after 2 h electrolysis detected by UV/Vis and 1H NMR spectroscopy

![](_page_11_Figure_1.jpeg)

**Fig. S30 a** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after 2 h electrolysis at -0.6 V in CO<sub>2</sub> saturated electrolyte, **b** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after 2 h electrolysis at -0.6 V in N<sub>2</sub> saturated electrolyte, **c** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after continuously supplying N<sub>2</sub> and CO<sub>2</sub> for 2 h without applied voltage, **d** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after 2 h electrolysis at open-circuit potential under ambient conditions, **e** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after 2 h electrolysis at open-circuit potential under ambient conditions, **e** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after 2 h electrolysis at open-circuit potential under ambient conditions, **e** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after 2 h electrolysis at open-circuit potential under ambient conditions, **e** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after 2 h electrolysis at open-circuit potential under ambient conditions, **e** UV-Vis absorption spectra of the 0.1 M KHCO<sub>3</sub> electrolyte stained with urea color agent before and after 2 h electrolysis at bare carbon cloth

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![](_page_12_Figure_1.jpeg)

Fig. S31 Under -0.6 V versus RHE for  $Cu^{II}$ -HHTP and  $Cu^{III}$ -HHTP during the electroreduction of N<sub>2</sub> and CO<sub>2</sub> processes

![](_page_12_Figure_3.jpeg)

**Fig. S32** The reaction pathway of \*NCONH\* formation over Cu<sup>III</sup>-HHTP. The structures of the initial (IS), transition (TS) and final (FS) states

![](_page_12_Figure_5.jpeg)

Fig. S33 In situ Cu K $\beta$  X-ray emission spectroscopy (XES) of a Cu<sup>III</sup>-HHTP and b Cu<sup>II</sup>-HHTP under different potentials

![](_page_13_Figure_1.jpeg)

**Fig. S34** The linear combination fitting (LCF) result of the Cu K-edge XANES spectra during -0.6V vs. RHE. **a** 0.5 h, **b** 1 h, **c** 1.5 h and **d** 2 h

![](_page_13_Figure_3.jpeg)

Fig. S35 The EXAFS WT signatures of a Cu foil, b CuO, c Cu<sup>II</sup>-HHTP, and d Cu<sup>III</sup>-HHTP.

For Wavelet Transform analysis, the  $\chi(k)$  exported from Athena was imported into the Hama Fortran code. The parameters were listed as follow: R range, 1-6 Å, k range, 0-12 Å<sup>-1</sup>; k weight, 2; and Morlet function with  $\kappa$ =10,  $\sigma$ =1 was used as the mother wavelet to provide the overall distribution

![](_page_14_Figure_1.jpeg)

![](_page_14_Figure_3.jpeg)

Fig. S37 In situ Raman spectra of Cu<sup>III</sup>-HHTP during the electrocoupling of N<sub>2</sub> and CO<sub>2</sub> at various potentials

![](_page_14_Figure_5.jpeg)

Fig. S38 Illustration of the different activation modes of  $N_2$  interacting with  $Cu^{II}$  active site for Cu<sup>II</sup>-HHTP

![](_page_15_Picture_1.jpeg)

Fig. S39 Electron density isosurface of  $CO_2$  molecule (left), CO(middle) and  $N_2$  molecule (right), the color bar represents the electrostatic potential scale

The electron density theoretical simulation analysis (Fig. S39) reveals that an electron-rich N atom exhibits in  $N_2$  molecule and an electron-deficient C atom in CO<sub>2</sub>/CO. Similarly, the charge accumulation area occurs in the Cu<sup>III</sup> part of Cu<sup>III</sup>-HHTP and the charge depletion one is found in the O part owing to the electron transfer from Cu to O, which can also be proofed by its XPS spectra (Fig. S4). As a result, Cu<sup>III</sup> and the adjacent O site in Cu<sup>III</sup>-HHTP possibly acts as the active centers towards the activation and coupling of  $N_2$  and CO<sub>2</sub> due to the electronic interaction.

![](_page_15_Figure_4.jpeg)

Fig. S40 DOS of a-b Cu<sup>II</sup>-HHTP and c-d Cu<sup>III</sup>-HHTP

Sample	Cu <sup>II</sup> -HHTP	Cu <sup>III</sup> -HHTP
θ	-167.18	-23
С	2.5	0.52
$\mu_{eff}$	4.47	2.04

**Table S1** The Weiss constant ( $\theta$ ), the Curie point (C) and effective paramagnetic moment (µeff) of pristine Cu<sup>II</sup>-HHTP and Cu<sup>III</sup>-HHTP catalysts

**Table S2** Comparison of the electrocatalytic urea production activity of Cu<sup>III</sup>-HHTP with previously reported urea electrosynthesis catalysts

Catalysts	Reactant	Electrolyte	Potential	FE	Urea yield	Refs.
Bi-BiVO <sub>4</sub>	N <sub>2</sub> , CO <sub>2</sub>	0.1 M KHCO <sub>3</sub>	-0.4	12.55	5.91	[S1]
PdCu	N <sub>2</sub> , CO <sub>2</sub>	0.1 M KHCO <sub>3</sub>	-0.4	8.92	3.36	[S2]
BiFeO <sub>3</sub> /BiVO <sub>4</sub>	N <sub>2</sub> , CO <sub>2</sub>	0.1 M KHCO <sub>3</sub>	-0.4	17.18	4.94	[S3]
Ni <sub>3</sub> (BO <sub>3</sub> ) <sub>2</sub>	N <sub>2</sub> , CO <sub>2</sub>	0.1 M KHCO <sub>3</sub>	-0.5	20.36	9.70	[S4]
InOOH	N <sub>2</sub> , CO <sub>2</sub>	0.1 M KHCO <sub>3</sub>	-0.4	20.97	6.85	[S5]
ZnO-V	$NO_2^-$ ,	0.2 M	-0.79	23.26	5.52	[S6]
Cu-TiO <sub>2</sub> -V <sub>0</sub>	$NO_2^-$ ,	0.2 M KHCO <sub>3</sub>	-0.4	43.10	4.16	[S7]
Zn nanobelts	NO, $CO_2$	0.2 M KHCO <sub>3</sub>	-0.92	11.26	15.13	[ <b>S</b> 8]
Cu <sup>III</sup> -HHTP	N <sub>2</sub> , CO <sub>2</sub>	0.1 M KHCO <sub>3</sub>	-0.6	23.09	7.78	This work

**Table S3** The ratios of  $Cu^{3+}$  and  $Cu^{2+}$  estimated by linear combination fitting (LCF) in the samples during -0.6V vs. RHE for ever-increasing eletrocatalysis time

	Fresh	0.5h	1.0h	1.5h	2.0h
Cu <sup>3+</sup>	100%	86%	74%	71%	69%
Cu <sup>2+</sup>	0%	14%	26%	29%	31%

Table S4 EXAFS fitting parameters at the Cu K-edge for various samples

Sample	Shell	N <sup>a</sup>	R (Å) <sup>b</sup>	σ2 (Å2·10 <sup>-3</sup> ) <sup>c</sup>	ΔE0 (eV) <sup>d</sup>	R factor (%)
Foil	Cu-Cu	12*	2.541-/+0.004	8.64-/+0.57	3.88-/+0.71	0.4
CuO	Cu-O	4*	1.951-/+0.012	5.19-/+0.81	-1.80-/+1.57	1.6
Cu <sup>II</sup> -HHTP	Cu-O4	4.2	1.961-/+0.021	4.81-/+2.23	-0.91-/+2.79	1.1
Cu <sup>III</sup> -HHTP	Cu-O4	4.0	1.963-/+0.006	6.25-/+0.41	0.24-/+0.88	0.6

<sup>*a*</sup> N: coordination numbers; <sup>*b*</sup> R: bond distance; <sup>*c*</sup>  $\sigma^2$ : Debye-Waller factors; <sup>*d*</sup>  $\Delta E_0$ : the inner potential correction. *R* factor: goodness of fit.  $S_0^2$  was set as 0.95 for Cu data, which was obtained from the experimental EXAFS fit of Cu foil reference by fixing CN as the known crystallographic value and was fixed to all the samples.

 Table S5 The detailed structures of DFT calculations

1. PATH-way		
1) Cu <sup>III</sup> -HHTP CO+N <sub>2</sub>		
CO+N <sub>2</sub>		
1.000000000000000		
21.879899978600001	0.0000000000000000000000000000000000000	000000000000000000000000000000000000000
-10.93495549100000	02 18.9466972881999	993 0.0000000000000000
0.00000000000000000	0 0.000000000000000	00 20.0552997588999986
C O Cu H N		
37 14 3 13 2		
Direct		
0.2351540743308887	0.5937259600732726	0.5318127880943928
0.5002144994936085	0.3739048054629709	0.5195145916775885
0.3129638712883224	0.5402472860964054	0.5252131160353037
0.3638237605059209	0.6628906264106521	0.5578717966820547
0.5719396027588837	0.1228312271487815	0.5249047307915046
0.4219569579245038	0.7917172155884087	0.5709607210721983
0.2965398144092964	0.7258554417234470	0.5365980003131967
0.8315707855603289	0.4496748639932521	0.5253399218417729
0.1666136519679294	0.6555142579299003	0.5270908686980522
0.7027316661452344	0.3839221063322057	0.5239000929182336
0.4351532181283101	0.6060864519411532	0.5606043162101562
0.6312364550668377	0.4421468740264676	0.5390597410469607
0 5711043014392110	0 3161395343314198	0 5081999726587817
0.3607711206544517	0.8542836578077294	0 5309966985166676
0.5115217763127725	0.1876929125265773	0 4907545039950632
0.6381189455020050	0.2511564637716162	0.5185900895612109
0.106832/19298/299	0.5257374601819832	0.5238258634656390
0.100052+17270+277	0.3102207374001017032	0.5180282323418208
0.7003307279070020	0.2512203040233808	0.5033720729955213
0.3783815057878747	0.5/171/61323307/7	0.5345094574549754
0.5765615057676742	0.31/83/318382/202	0.5034088830056203
0.5075558498718520	0.3140343103024202	0.50540888555550205
0.0337336319266739	0.5000270850204216	0.5236258024571857
0.103099414/94/104	0.3909270830294210	0.5250536924571657
0.7004200521175654	0.44/021100/3/2/44	0.5204207025505745
0.7030974033244340	0.5177063609007020	0.5200500574000088
0.4240320207600190	0.0331720019392021	0.5375110405105752
0.0538090745287730	0.10/0004439000421	0.5501114975259510
0.3013832883080348	0.1213403940834298	0.5574291058795455
0.50012084//549199	0.4413499856888479	0.53/85138403/8448
0.42//4038/0190650	0.0040549414475451	0.5704393624636571
0.3039072202953634	0.59/591/458238994	0.5380234415659356
0.8326626398/99610	0.3833190284291283	0.5211023528266481
0.1/308/1615051929	0.5291953892770201	0.5279649302458539
0.2322872991512927	0.658/37/832060832	0.5313348200200001
0.5086241122666340	0.1232766566730540	0.5018101616873735
0.2983003240025380	0./894660918836087	0.5230974155779510
0.5273294977453348	0.6000794048455977	0.6221450627325363
0.0471611637235062	0.4671112068080360	0.5214799570669711
0.3676859449391772	0.9149927723323320	0.5161525367480022
0.8921722998545191	0.5071845049227464	0.5269376754432544

0.4415820141699285 0.3711230984372968 0	5225142636656855
0.5644359682598636 0.0618372661029405 0.	5396401306976224
0.4510943073564760 0.0629197188983910 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510943073564760 0.4510940000000000000000000000000000000000	4960355018385739
0.5603430297671776 0.4944405073927500 0.	5539865482169731
0.8940405440220084 0.3881726135561498 0.	5189374498831502
0.4805025067669794 0.9160770437068760 0.	5639877387864383
0.3819787004640967 0.4859106819684099 0.	5198353687731734
0.0411689439195453 0.5842181605770149 0.	5215268449867159
0.5050622501369860 0.6189494045539097 0.	5669842809017479
0.5860677443663332 0.6302205857671213 0.4	6437177771813007
0.4932022636624138 0.4967326466925301 0.4	4400649127389439
0.4656755380957024 0.9890451840224346 0.	5289915907349004
0.9690291333575152 0.4864027490765626 0.	5216610301395405
0.4662608594915683 0.4807818397289799 0.	5292671760571392
0.4748171729698685 0.7124081914560473 0.:	5857556672826071
0.2522151635160618 0.7904622696395855 0.	5039388610617251
0.1731458156157589 0.4794248670556092 0.	5299675468857669
0.4584686071797027 0.2668425774676857 0.4	4916716559932002
0.6824433259369551 0.1850368095691964 0.	5457456066351831
0.7678398288662115 0.4982986437344562 0.	5278015014891239
0.6774849165656812 0.4916562687858060 0.	5536853316778868
0.7708453330877393 0.2708904932174004 0.	5142547512623210
0.4623987485456961 0.1853940443435232 0.4	4766252174010754
0.2698697720301614 0.4904658195758736 0.	5065046425857610
0.1626367604309701 0.7031779697745316 0.	5289674146880671
0.4709710166079474 0.7954138894473338 0.:	5877912165246620
0.4522209910174851 0.4861398907293072 0.4	4126718847359655
0.4709459472925703 0.5329459777334112 0.4	6545573971001654
0.4461488668238441 0.4781517339525270 0.4	6245622189569684
2) $Cu^{III}$ -HHTP $CO$ + $N_2H$	
CO+N <sub>2</sub> H	
1.00000000000000	
21.8798999786000010 0.0000000000000000	0.0000000000000000000000000000000000000
-10.934955491000002 18.946697288199999	03 0.0000000000000000000000000000000000
0.0000000000000 0.000000000000000000000	20.0552997588999986
C O Cu H N	
37 14 3 14 2	
Direct	
0.2339594895692932 0.5878723220754704 0.:	5229113656856982
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	0.4784145133565547	0.7127232039087683	0.5715494515139670
	0.2526866860252957	0.7905365521620205	0.5012512377300256
	0.1742795927187946	0.4797101002032915	0.5241196371134624
	0.4592175327279238	0.2638949000433186	0.5327480037368348
	0.6839275044502531	0.1839009245147776	0.5391374813615638
	0.7709580963953393	0.4995134821329911	0.5303843613010638
	0.6825690600665995	0.4981243169395179	0.5329080712143385
	0.7722622126319414	0.2714613653348159	0.5166754946286318
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	0.1644965165447008	0.7035586693307345	0.5307088500173678
	0.4753912571143919	0.7961479534157070	0.5721746547566444
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	0.4629548801849574	0.4979653698779082	0.6369593083235012
(	6) Cu <sup>III</sup> -HHTP NH <sub>2</sub> CO	$ONH_2$	
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	-10.93495549100000	02 18.9466972881999	993 0.0000000000000000
	0.00000000000000000	0 0.000000000000000	00 20.0552997588999986
	C O Cu H N		
	37 14 3 17 2		

Direct

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0.5700512855092422	0.1260519202654546	0.5214385099850258
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0.6349261090181951	0.4500675798299184	0.5338146581689009
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o) $Cu^{-}$ -HHIP $CO+N_2I$	7	

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 C O Cu H N
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<i>10) Cu<sup>11</sup>-HHTP NHCO</i>	NH (NH			
NHCONH				

1.00000000000000

21.8798999786000010 -10.934955491000002 18.9466972881999993 0.0000000000000000 0.00000000000000000 0.0000000000000 20.0552997588999986 C O Cu H N 37 13 3 14 2 Direct 0.2380109811898911 0.5909893560029359 0.5296174426467724 0.5074693628606946 0.3824981135337938 0.5038548118417713 0.3135961408151556 0.5348701501072840 0.5220994659573417 0.3683394031746751 0.6597785897071766 0.5503106505419320 0.5721535014164096 0.1245870411766850 0.5254805782401324 0.4283287847998857 0.7897400402362957 0.5614920227997909 0.3011440921272578 0.7236564410315123 0.5341415800763863 0.8330942754494952 0.4509579646627338 0.5272893268065514 0.1707378993610442 0.6543393512495549 0.5306158397139988 0.7043011539472922 0.3854182568651506 0.5245104978739280 0.4384823851969541 0.6013242507901220 0.5511262677766939 0.6317253463680480 0.4416805842565408 0.5451866209899537 0.5741787853619632 0.3202652585465616 0.5008824234890613 0.3651109344768224 0.8525105100561541 0.5296126158425855 0.5112803632395028 0.1898887758071629 0.4933987272273149 0.6388071864276256 0.2530609759111724 0.5170647112328991 0.1091402850566032 0.5245056222801141 0.5251523401023513 0.7695415300131621 0.3203819751739487 0.5181715978214064 0.5741152287516778 0.2537357672052874 0.5012555568762430 0.3792308560773734 0.5359897010366409 0.5280765297596789 0.5138121737398434 0.3233635030761680 0.4854632090950601 0.6372940271890803 0.3838769267357278 0.5244803132073343 0.1072094477218806 0.5901862373939933 0.5280901904326023 0.7679471032309881 0.4492904641519638 0.5280677024134566 0.7051356126606840 0.3191363617640397 0.5195717992684197 0.4295779521064487 0.8535805718409502 0.5514384737937602 0.6360596216446444 0.1888263805328539 0.5293981316761448 0.3669166888200585 0.7248659552737571 0.5505126826080743 0.5681762513799006 0.4426261683054896 0.5370102034195827 0.4323374884783595 0.6600488191888658 0.5612196812453409 0.3065288332011964 0.5938332841801027 0.5333549915019211 0.8338697041903124 0.3843525296677422 0.5219395234103302 0.1750339864873943 0.5269799826135145 0.5266081624266539 0.2359956449620819 0.6566042576452575 0.5309670308959752 0.5082240420192171 0.1249880525006241 0.5048485202949463 0.3024476863382357 0.7874107967904952 0.5228371921905968 0.4902903117399152 0.5563263738085382 0.6699221788300387 0.0487454788038040 0.4665830208796462 0.5225173029827298 0.3705339132613975 0.9132324154092503 0.5175283904083028 0.8939661966467394 0.5080548532238472 0.5298666113687898 0.4508945113156910 0.3856622560785415 0.4984816280037799 0.5651243696617751 0.0635823286092941 0.5398829159825302 0.4507503442966003 0.0643475154623288 0.5009808006973887 0.5592056620945823 0.4938103885721920 0.5552826107529754 0.8952913788975158 0.3892247571723727 0.5201280816468773

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NHCONH<sub>2</sub>
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                      0.0000000000000 20.0552997588999986
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Direct
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-10.93495549100000	)2 18.9466972881999	993 0.0000000000000000
0.0000000000000000000000000000000000000		00 20 0552997588999986
$C \cap Cu \in N$		2010202///2007///00
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Direct		
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0.1713375132375576	0.5276798893018597	0.5199197939494997
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2. $CO_2$ reduction reaction
1) Cum-HHTP CO2
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0.0000000000000 0.000000000000 20.0552997588999986
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0.3682364154919519 0.6560637214067674 0.5390234507145039
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0.3031302207238834 0.7234746728448611 0.5300328523309794
0.8325685722877857 0.4541879532608354 0.5287464869949169
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C O Cu H
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5) <i>Cu<sup>II</sup>_HHTP COOH</i>	0.4437314200430423	0.3873327078370278
1 00000000000000		
21 87080007860000		
-10.03/055/0100000	$\frac{10}{12}  \frac{18}{18} \ \frac{9}{66072881000}$	993 0.0000000000000000000000000000000000
	02  10.9400972001999	00 20 0552997588999986
$C \cap Cu H$	0 0.00000000000000000000000000000000000	00 20.0552777500777700
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